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Docket No.: 3313-1056P

Application No. 10/712,045
Amendment dated September 5, 2006
Reply to Office Action of May 5, 2006

REMARKS

Claims 1-8, 10-19, 21 and 22 are now present in this application.

The specification, abstract, and claims 1, 4, 5, 7, 8, 12, 15, 16 and 18 have been amended, and claims 9 and 20 have been cancelled without prejudice or disclaimer of the subject matter contained therein. Reconsideration of the application, as amended, is respectfully requested.

Amendments to the Specification

Certain changes have been made to the specification to correct grammatical and typographical errors. It is also noted that a brief description of an additional embodiment has been provided. Because support for these changes can be found in the originally filed specification, claims and drawings, it is respectfully submitted that no new matter is present.

Objection to the Claims

Claims 9 and 20 stand objected to under 37 CFR 1.75(c). In view of the foregoing amendments, in which claims 9 and 20 have been cancelled, and certain portions thereof have been incorporated into independent claims 1 and 12, it is respectfully submitted that this objection has been addressed. Reconsideration and withdrawal of any objection to the claims are respectfully requested.

Rejection under 35 USC 112

Claims 9 and 20 stand rejected under 35 USC 112, second paragraph. This rejection is respectfully traversed.

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In view of the foregoing amendments, in which claims 9 and 20 have been cancelled, and certain portions thereof have been incorporated into independent claims 1 and 12, it is respectfully submitted that all claims particularly point out and distinctly claim the subject matter of the instant invention. Accordingly, reconsideration and withdrawal of the 35 USC 112, second paragraph rejection are respectfully requested.

Rejection under 35 USC 102

Claims 1-2, 6-8, 10 and 11 stand rejected under 35 USC 102(b) as being anticipated by Naarmann et al., U.S. Patent 5,174,867. This rejection is respectfully traversed.

Claims 1-8, 10-16, 18, 19, 21 and 22 stand rejected under 35 USC 102(b) as being anticipated by Lessner et al., U.S. Patent 6,001,281.

Naarmann

Although Naarmann uses imidazole to synthesize conducting polymer, the composition and method utilized by Naarmann differs from that of the present application.

In Naarmann, imidazole is copolymerized with pyrrole monomer by electrochemical polymerization (see column 2, lines 64-67; column 3, lines 40-44 and 55-65; and column 4, lines 54-63). Imidazole is utilized in Naarmann as a monomer but not a retardant (see column 2, lines 6-12). In addition, the formula in Naarmann does not contain Fe(III) ion, and therefore imidazole will not complex with Fe(III) ion to stabilize the precursor solution. In the present application, however, polymeric electrolyte is synthesized by chemical oxidative polymerization but not electrochemical polymerization. Therefore, imidazole does not act as a monomer, but instead acts as a retardant to complex with an oxidant-Fe(III) ion to stabilize the oxidizing agent.

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Lessner

Lessner et al. discloses a completely different compound and method from the present application. In Lessner, low-boiling compounds are used with oxygen atoms such as cyclic ethers and ketons, in order to weakly complex with Fe(III) ion to stabilize the diluted oxidizing agent (see column 4, lines 22-26, 30-31, 40-42 and 58-61, and column 5, lines 3-10). In the present application, however, a thick oxidant solution is added with the compound of $\text{—}\overset{\text{I}}{\text{C}}=\text{N—}$ functional group as a strong retardant to fabricate capacitors.

Although Lessner et al. uses pyridine as a retardant to stabilize the precursor solution for synthesizing conducting polymer, the properties were not gratifying (see example 1E). Therefore, in the disclosed example, the formula containing pyridine was merely used as a comparison to present the effect of the idea. In the present application, however, the formula containing pyridine can successfully converse into the polymeric electrolyte of a solid electrolytic capacitor with extraordinary properties. Obviously, Lessner's idea, in which a low-boiling compound acts as a retardant in the precursor solution, is not suitable for the case with pyridine, so only compounds with oxygen atoms such as cyclic ethers and dioxanes are claimed as a retardant. Pyridine was *excluded* from Lessner (see column 10, lines 32-36). In the present application, the above-noted problem has been overcome, and a solid electrolytic capacitor with extraordinary properties is fabricated utilizing a simpler procedure than that of Lessner.

Rejection under 35 USC 103

Claim 17 stands rejected under 35 USC 103 as being unpatentable over Lessner et al.

This rejection is respectfully traversed.

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The Examiner asserts that the prior art "discloses the addition of non-volatile organic base such as imidazole to inhibit the polymerization of the monomer mixture (Col-3, Ln 3-5)," and that "such an addition is well known in the art at the time of the disclosure of the invention by the applicants (See, Matsumoto et al., JP 06-112094, Abstract)."

The prior art (EP0615256B1) quoted in Lessner focuses on the coating of conducting polymer on a substrate, although a non-volatile organic base such as imidazole is added into the formula to inhibit the polymerization of the monomer mixture (see column 1, lines 1-10, and column 6, lines 36-37). Thus the goals of the foresaid prior art and the present application are different. In addition, Lessner et al. considers that the organic base in the precursor will be left in the pores of a capacitor and therefore cause failure in the manufacture of solid electrolytic capacitors (see column 3, lines 5-7). However, the present application has successfully used such non-volatile organic bases to fabricate excellent solid electrolytic capacitors with conducting polymer as electrolyte. Therefore, the present application has achieved an end result previously thought impossible.

Moreover, in comparing the present application with Matsumoto, it is noted that the two cases are not related. According to Matsumoto, pyridine-ring compounds are added into the non-transition metal oxidizing agent such as ammonium persulfate, or monomer solution, for promoting the impendence of capacitors (see examples 1-8). It is noted, however, that pyridine-ring compounds will *not* complex with the non-transition metal oxidizing agent or monomer for stabilizing the precursor solution, and therefore an oxidant solution with high concentration cannot be used and the procedure for making a solid electrolytic capacitor set forth in Matsumoto remains extremely complicated. In the present application, however, pyridine will complex with

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Fe(III) oxidant to reduce the oxidizing capability of the Fe(III) oxidant at an ambient temperature, so an oxidant solution with high concentration can be used to simplify the manufacture of solid electrolytic capacitors. Also, the properties of such solid electrolytic capacitors are improved over the prior art. Therefore, the role of pyridine ring compounds in Matsumoto is not comparable to the present application.

Although pyridine and imidazole may have been used in some cases for synthesizing conducting polymer, before the present application, Fe(III) oxidizing agents had not been successfully used with these compounds for fabricating solid electrolytic capacitors with satisfactory properties. The results of Lessner (see Example 1E) demonstrates that fabricating excellent solid electrolytic capacitors with these compounds has been beyond the reach of one of ordinary skill in the art, based on common sense. In the present application, the traits of compounds of a $\text{—}\overset{\text{I}}{\text{C}}=\text{N—}$ functional group, combined with the novel concept of an oxidant solution with high concentration, simplifies procedure and produces excellent solid electrolytic capacitors. The solid electrolytic capacitors of the present application demonstrate higher capacitance and superior impedance to the solid electrolytic capacitors made by the prior art.

It is therefore respectfully submitted that the prior art utilized by the Examiner fails to teach or suggest the mixture solution and method of independent claims 1 and 12 of the present application, as well as their dependent claims. Reconsideration and withdrawal of the 35 USC 102 and 103 rejections are therefore respectfully requested.

Conclusion

Favorable reconsideration and an early Notice of Allowance are earnestly solicited.

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In the event that any outstanding matters remain in this application, the Examiner is invited to contact the undersigned at (703) 205-8000 in the Washington, D.C. area.

Pursuant to 37 C.F.R. §§ 1.17 and 1.136(a), the Applicants respectfully petition for a one (1) month extension of time for filing a response in connection with the present application and the required fee of \$120.00 is authorized to be charged to Deposit Account No. 02-2448.

If necessary, the Commissioner is hereby authorized in this, concurrent, and future replies, to charge payment or credit any overpayment to Deposit Account No. 02-2448 for any additional fees required under 37 C.F.R. §§ 1.16 or 1.17; particularly, extension of time fees.

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**MIXTURE SOLUTION ~~SOLUTION~~ FOR PREPARING A CONDUCTIVE
POLYMER TO PRODUCE ~~SOLID~~ SOLID ELECTROLYTIC CAPACITORS AND
ITS ~~MEYHOD~~ METHOD**

This nonprovisional application claims priority under 35 U.S.C. § 119(a) ~~on~~ of
Patent Application No(s) ~~No.~~ 91136357 filed in ~~TAIWAN~~ Taiwan, R.O.C. on December
17, 2002, the entire contents of which are hereby incorporated by reference.

5 **BACKGROUND ~~BACKGROUND~~ OF THE INVENTION**

Field of Invention

The invention relates to a mixture solution for preparing a conductive polymer to
produce a solid-state electrolytic capacitor. The conducting polymer is formed from the
10 mixture of monomer and oxidant solution. Particularly, the oxidant solution is high
concentration, and is also added a five or six-member ring compound with a functional
group $-\overset{\overset{|}{\text{C}}}{\text{C}}=\text{N}-$, which acts as a retardant for the polymerization.

Related Art

15 For a long time, the development of electrolytic capacitors has been focusing on
improving the conductivity of electrolyte so as to reduce the equivalent series resistance
(ESR) and impedance and improve the reliability of capacitors. Because conductive
polymer has higher conductivity than liquid-state electrolyte or solid-state organic
semiconductor complex salt (such as TCNQ complex salt) and it becomes insulator in
20 higher temperature, it is perfectly used as the electrolyte of electrolytic capacitors.

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In ~~United States patent~~ U.S. Patent No. 4803596, Gerhard Hellwig, Stegen et al. first disclosed that conductive polymer would be used as the electrolyte of an electrolytic capacitor. After the positive foil of an electrolytic capacitor is dropped with the monomer and oxidant solution, respectively, the monomer polymerizes by the oxidant under proper
5 condition. By this method, however, the conductive polymer monomer is not fully mixed with the oxidant, so the reaction is not uniform.

Therefore the method for producing conductive polymer electrolytes of solid-state electrolytic capacitors is improved by mixing conductive polymer monomer with dopant, oxidant and solvent, and then making the capacitor element impregnated with such the
10 mixture solution. Finally the monomer absorbed in the element converts into conductive polymer. If the reactivity of the oxidant in the mixture is not inhibited, however, the conductive polymer monomer polymerizes as soon as it mixes with the oxidant. As a result, the mixture solution is unable to permeate into the pores of the capacitor element.

Friedrich Jonas et al. disclosed in ~~United States~~ U.S. Patent No. 4910645 a series of
15 specific polythiophenes can apply to the electrolyte of solid state electrolytic capacitors. The method is simply to dip the capacitor element in the mixture of thiophene monomer and oxidant solution without polymerization retardant and then thiophene monomer polymerize at high temperature. Unfortunately, the stability of the mixture of the thiophene monomer and oxidant solution at room temperature decreases with the high concentration
20 both of the monomer and oxidant. Therefore the ~~preposal~~ proposed method uses a large amount of solvent to dilute the monomer and oxidant. As a consequence, when a capacitor element is impregnated with the mixture solution, it only obtains a little conductive polymer, and the rest is the remaining reagent and a large amount of solvent. Multiple steps of

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impregnation and polymerization processes are required by this method in order to produce enough conductive polymer to fill the space in a capacitor element.

Phpip M. Lessner et al. disclosed in ~~United States~~ U.S. Patent No. 6,056,899 a process that uses a kind of cyclic ethers (such as tetrahydrofuran) to mix with an oxidant of ferric ion to produce a complex, reducing the oxidation ability of the oxidant so that the mixture solution of the conductive polymer monomers and oxidant is kept stable for a long period of time. After the capacitor element is impregnated with the mixture solution, the cyclic ether is evaporated at a higher temperature to make the oxidant ~~induces~~ induce a polymerization of conductive polymer. Because the ~~said~~ cyclic ethers disclosed in the invention (such as tetrahydrofuran) ~~has~~ have little ability as a polymerization retarder to ~~stabilize~~ stabilize the mixture solution of monomer and oxidant, a large amount of such cyclic ether is consumed in order for it to ~~stabilize~~ stabilize the mixture solution and, at the same time, dilutes the mixture solution of conductive polymer monomer and oxidant. As a result, a capacitor element still needs to be impregnated and polymerized many times to acquire the a sufficient amount. Thus, ~~it this patent~~ still has the shortcomings of ~~fabreating~~ fabricating a solid-state electrolytic capacitors by a complicated process.

SUMMARY OF THE INVENTION

To solve the aforesaid problems, the present invention provides a simple method for forming conductive polymer as the electrolyte of a solid-state electrolytic capacitor. The main object is to use the simple process steps of the impregnation and polymerization to obtain conducting polymer as the electrolyte of a solid electrolytic capacitor. By this invention, a solid-state electrolytic capacitor exhibiting excellent capacitor characteristics is acquired.

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The invention provides a mixture solution and its method for preparing the conductive polymer to produce a solid-state electrolytic capacitor. The method is that polymerization retarder is added into a high concentration oxidant solution in advance, to reduce the oxidizing ability of the oxidant so that polymerization does not occur until the temperature is higher than that of impregnation. Specifically, polymerization retarder is a kind of five- or six-member ring compound with a functional group $\text{—}\overset{\text{I}}{\text{C}}=\text{N—}$. The said five- or six- member ring compounds having a functional group $\text{—}\overset{\text{I}}{\text{C}}=\text{N—}$ are more effective in stabilizing the mixture solution of monomer and oxidant than the said cyclic ethers. So accordingly, the amount of the said five- or six-member ring compounds having a functional group $\text{—}\overset{\text{I}}{\text{C}}=\text{N—}$ in the oxidant solution is minor and will not obviously dilute the mixture solution. In addition, the solvent used in this oxidant solution, such as methanol, is high solubility highly soluble to the oxidant containing ferric ion, so that the amount of the solvent in the mixture solution can be limited in an a minor proportionation. As a consequence consequence, the mixture of conducting polymer monomer and oxidant solution has good stability at room temperature, and there is sufficient time to impregnate a capacitor element with the mixture solution. Finally, the majority of the absorbed mixture solution in the impregnated capacitor element converts into conducting polymer, because the retarder retardant and solvent in the mixture solution are in a tiny minority. Therefore, the steps of the impregnation and polymerization of conducting polymer as the electrolyte of a solid electrolytic capacitor are reduced.

Further scope of the applicability of the present invention will become apparent from the detailed description given hereinafter. However, it should be understood that the detailed description and specific examples, while indicating preferred embodiments of the

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invention, are given by way of illustration only, since various changes and modifications within the spirit and scope of the invention will become apparent to those skilled in the art from this detailed description

5 **BRIEF DESCRIPTION DESCRIPTION OF THE DRAWINGS**

The present invention will become more fully understood from the detailed description given hereinbelow and the accompanying drawings which are given by way of illustration only, and thus are not limitative of the present invention, and wherein:

Fig. 1 is the illustration of the polymerization of 3, 4-ethylenedioxythiophene,
10 chemically oxidized by Fe(III) tosylate;

Fig. 2 is a figure of the specific viscosity versus time for the mixture solution of 3, 4-ethylenedioxythiophene and various oxidants at room temperature;

Fig. 3 is a composition (wt) table of different mixture ~~solutions~~solutions; and

Fig. 4 is a table illustrating capacitor characteristics of different embodiments of the
15 invention .

DESCRIPTION OF THE PREFERRED EMBODIMENT

Polyaniline, polypyrrole, polythiophene and their derivatives, such as poly(3, 4-ethylenedioxythiophene) are common commercial conductive polymers and are widely
20 used as the electrolyte of solid electrolytic capacitors-widely. Fig. 1 illustrates the chemical polymerization of 3, 4-ethylenedioxythiophene (EDT) that uses Fe(III) tosylate as oxidant, which oxidizes EDT to produce poly (3, 4-ethylenedioxythiophene) and toluenesulfonic acid at the same time. Toluenesulfonic acid dopes poly (3, 4-ethylenedioxythiophene).

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Fe(III) tosylate as an oxidant in this invention, like other ferric oxidants, is solid and is proper to be used after dissolved in solvent. In the invention, the solvent that dissolves the oxidant may be alcohol, acetone, water, or any mixtures containing any of these. To reduce the proportion of solvent in the mixture solution of monomers and oxidant solution, the best
5 result is obtained when the concentration of oxidant reaches over 50wt% in the mixture solution.

In addition to Fe(III) tosylate, ferric sulfonic acid, ferric perchloride and ferric chloride, the residual of which itself may act as a dopant, ~~but, and other oxidants, for example—~~ such as ammonium persulfate, need dopants to make poly 3,
10 4-ethylenedioxythiophene exhibit high conductivity. The usual dopants include AsF_6^- , BF_4^- , PF_6^- , I and single proton or multi-proton acid, such as sulfuric acid, phosphoric acid, citric acid, HClO_4 , perchloric acid, and relevant ~~derivative~~ derivatives or salts, e.g. polystyrene sulfonic acid (PSS).

Besides relating to the oxidation potential of the monomer, the oxidation
15 polymerizing reaction of the conductive polymer monomer is also affected by the oxidizing capability of the oxidant. For the oxidant containing ferric ion, the oxidizing capability of the ferric ion may decay when it forms a complex with specific molecules. Therefore, the stronger the complex bonding is, the lesser the oxidizing capability of the oxidant.

The invention uses a compound of a five-member or six-member ring with a
20 functional group $-\overset{\text{I}}{\text{C}}=\text{N}-$, which produces complex bonding with ferric ions of the oxidant to reduce the oxidizing capability of the oxidant with ferric ion at room temperature. As a result, the mixture of monomer and oxidant is stable at room temperature, even if the amount of the solvent in the mixture is tiny. The oxidant which forms a complex with the said polymerization retardant can still oxidize the monomer to produce conducting polymer

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in a high temperature environment. The compound of a five-member ring with functional group $-\overset{|}{\text{C}}=\text{N}-$ can form stronger complex bonding with ferric ion than that of oxygen-containing compounds. Therefore, only a small amount of the compound of a five-member ring with a functional group $-\overset{|}{\text{C}}=\text{N}-$ has ~~a~~ ~~has~~ higher efficiency to retard the oxidizing capability of the oxidant containing ferric ion at room temperature. For the oxidant containing ferric ion, its complexes with a six-member ring with a functional group $-\overset{|}{\text{C}}=\text{N}-$ has ~~a~~ the strongest complex bonding. The mixture solution of the monomer and oxidant is made by mixing the high concentration oxidant with a small amount of polymerizing retardant, which has a five-member or six-member ring with a functional group $-\overset{|}{\text{C}}=\text{N}-$, for which the monomer and oxidant are main components. The element of the capacitor not only has adequate impregnating time but also acquires conductive polyelectrolytes with sufficient thickness in the limited impregnating times. Thus conductive polymer electrolytes of solid-state electrolytic capacitors with excellent performance are produced.

The object of the invention is to add a polymerization retardant with a five-member or six-member ring compound with a functional group $-\overset{|}{\text{C}}=\text{N}-$ in the high concentration oxidant solution to prevent the monomer from immediately reacting when mixing the oxidant solution with the conductive polymer monomer at room temperature. A capacitor element is then impregnated in ~~such~~ the mixture solution, and conductive polymer forms at a higher temperature later. In such a way, the high concentration oxidant solution can be used, eliminating the disadvantage that the capacitor needs ~~multi~~ multiple impregnation and polymerization processes to form electrolytes of adequate thickness. The method also offers adequate impregnation time for capacitor elements and produces conductive polymer electrolytes that deeply penetrate into the micropores of the element foil. Thus, capacitor

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elements acquire polymer electrolytes of adequate thickness and excellent performance after being impregnated with the mixture solution and making a polymerization in only a few times.

At room temperature or slightly higher, the invention can properly reduce the oxidizing capability of the oxidant containing ferric ion. Five-member ring compounds (with functional group $\text{—}\overset{\text{I}}{\text{C}}=\text{N—}$), such as imidazole, 2-methylimidazole, pyrazole, triazole, pyridine and pyridazine, as well as their derivatives, may stabilize the mixture solution of the conductive polymer monomer and oxidant which acts as the polymerizing retardant for conductive polymer monomer. Since the complex capability of imidazole, 2-methylimidazole, pyrazole, triazole and their derivatives is much greater with ferric ~~ion~~ ion than that of compounds containing oxygen, only a little small amount of any one of imidazole, 2-methylimidazole, pyrazole, triazole or their derivatives is required to reduce the oxidizing capability of ferric oxidant at room temperature. By adding the six-member ring compound containing a functional group $\text{—}\overset{\text{I}}{\text{C}}=\text{N—}$, which has stronger complex capability with ferrous oxidant, such as pyridine, pyridazine or 1,2-diazine and their derivatives, conductive polymer electrolytes with high electrical conductivity and high stability are also produced, ~~too~~. The mixture solution of monomer and oxidant (these are the main components) is made by mixing high concentration oxidant with a little polymerization retardant so that the capacitor element not only has enough time for impregnation, but also acquires an adequate thickness of conductive polymer electrolytes. Thus, conductive polymer of solid-state electrolytic capacitors ~~of~~ with excellent performance are produced.

The following is a detailed description of the technology required for achieving the objectives of the invention:

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Embodiment 1

In order to compare the long- term stability of the mixture solutions containing different retardants, the six different mixture solutions of EDT monomer, Fe(III) tosylate solution and different polymerization retardants are made according to ~~the table~~ Table 1.

- 5 The changes of specific viscosity of these six mixture solutions over time are measured at room temperature (25°C). The results are listed in Fig. 3.

The results in Fig. 3 show that the mixture solution of EDT monomer and Fe(III) tosylate solution without any polymerization retardant is most unstable in these six mixture solutions. The specific viscosity of such the mixture solution increases quickly, showing
 10 that the polymerization is occurring rapidly, and the capacitor element does not have enough time for impregnation with such ~~the an~~ unstable mixture solution. Although the oxidant solution containing tetrahydrofuran (THF) can slightly delay polymerization of monomers in the mixture solution, the amount of tetrahydrofuran which is required is much more than the four retardants above—: imidazole, 2-methylimidazole, triazole and pyridine.
 15 ~~Reversely~~Conversely, the variations of specific viscosity of all the mixture solutions containing the reductants of five- or six-member ring with the functional group $—\overset{|}{C}=N—$ are not very obvious during the measuring period of time, which shows that such the mixture solutions exhibit better ~~satability~~ stability at room temperature and a capacitor can be impregnated sufficiently in such ~~mixture-mixture~~ solutions.

20

Embodiment 2

An anode aluminum foil, a ~~ea~~the~~ed~~cathode aluminum foil and Manila paper are wound together to prepare ~~an a~~ capacitor element. The element is impregnated in a mixture solution composed of EDT monomer, Fe(III) tosylat solution and imidazole. The
 25 proportions of the ingredients in the mixture solution are listed in Table 1. The capacitor element sucking the mixture solution is heated at 100°C for 10 minutes. After two times of

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impregnation and polymerization, the capacitor element is cleaned up with methanol and dried at 100°C. Then the capacitor element is set into an aluminum metallic case and sealed with a rubber member. After the aging process is completed, a winding type aluminum solid state electrolytic capacitor is produced. For the features of the capacitor according to this

5 Embodiment refer to Fig. 4.

Embodiment 3

In the same manner of Embodiment 2, an anode aluminum foil, a cathode aluminum foil and Manila paper are wound together to prepare an a capacitor element. The element is impregnated in a mixture solution composed of EDT monomer, Fe(III) tosylat solution and 2-methyl imidazole. The proportions of the ingredients in the mixture solution are listed in Table 1. The capacitor element sucking the mixture solution is heated at 100°C for 10 minutes. After two times of impregnation and polymerization, the capacitor element is cleaned up with methanol and dried at 100°C. Then the capacitor element is set into an aluminum metallic case and sealed with a rubber member. After the aging process is completed, a winding type aluminum solid state electrolytic capacitor is produced. For the features of the capacitor according to this Embodiment refer to Fig. 4.

Embodiment 4

20 In the same manner of Embodiment 2, an anode aluminum foil, a cathode aluminum foil and Manila paper are wound together to prepare a capacitor element. The element is impregnated in a mixture solution composed of EDT monomer, Fe(III) tosylat solution and imidazole. The proportions of the ingredients in the mixture solution are listed in Table 1. The capacitor element sucking the mixture solution is heated at 100°C for 10 minutes. After two
25 times of impregnation and polymerization, the capacitor element is cleaned up with

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methanol and dried at 100°C. Then the capacitor element is set into an aluminum metallic case and sealed with a rubber member. After the aging process is completed, a winding type aluminum solid state electrolytic capacitor is produced. For the features of the capacitor according to this Embodiment refer to Fig. 4.

5

Embodiment 5[[4]]

In the same manner of Embodiment 2, an anode aluminum foil, a ~~cathode~~ cathode aluminum foil and Manila paper are wound together to prepare a capacitor element. The element is impregnated in a mixture solution composed of EDT monomer, Fe(III) tosylate solution and pyridine. The proportions of the ingredients in the mixture solution are listed in Table 1. The capacitor element sucking the mixture solution is heated at 100°C for 10 minutes. After two times of impregnation and polymerization, the capacitor element is cleaned up with methanol and dried at 100°C. Then the capacitor element is set into an aluminum metallic case and sealed with a rubber member. After the aging process is completed, a winding type aluminum solid state electrolytic capacitor is produced. For the features of the capacitor according to this Embodiment, refer to Fig. 4.

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Comparative Embodiment 1

After the Fe(III) tosylate constituting the mixture solution listed in the second set of Fig. 2 is dissolved with methanol and mixed with tetrahydrofuran, it is mixed with EDT monomer. Finally, the mixture solution is impregnated with the capacitor element. The impregnating mixture solution sucked by the capacitor element reacts for 10 minutes at 100°C. After two times of impregnating and reaction, the capacitor element is cleaned up with methanol and dried at 100°C. The ~~capacitor~~ capacitor element is set into an aluminium

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metallic case and sealed with a rubber member. Finally, aging is carried out to mend the dielectric layer of the capacitor foil, and the capacitor is thus produced completely.

Refer to Fig. 4 for characteristics of the capacitor according to this comparative Embodiment. Because there is a great deal of tetrahydrofuran in the mixture solution, fewer
5 conductive ~~polymer-polymers~~ are to be aquired. Accordingly, with the same impregnating times, the conductive polymer electrolytes acquired by the capacitor element being immersed in this type of mixture solution ~~is-are~~ unable to fill the entire ~~sapee-space~~ of the capacitor element, which is obviously not as good as the embodiments described above.

10 Comparative Embodiment 2

Refer to the sixth set of mixture solution in Fig. 2 for the amout of Fe(III) tosylate, methanol and EDT monomer in this Embodiment, among which only the amount of pyridine is increased to 0.266g. This mixture solution is then put into the element formed according to the specifications of the second embodiment of the invention, and allowed to
15 react for 10 minutes at 100°C. Consequently, the mixture solution sucked by the element converts into brown gel, and, as such, the capacitor has no capacity. Because the amount of pyridine in this mixture solution is higher than in the Embodiment 4, the retard effect of pyridine on the oxidant is so much stronger that the EDT monomer in the mixture solution ~~can't~~ cannot convert into conductive polymer due to insufficient oxidizing of the oxidant.

20

~~While the preferred embodiments of the invention have been set forth for the purpose of disclosure, modifications of the disclosed embodiments of the invention as well as other embodiments thereof may occur to those skilled in the art. Accordingly, the appended claims are intended to cover all embodiments which do not depart from the spirit~~

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~~and scope of the invention.~~ The invention being thus described, it will be obvious that the same may be varied in many ways. Such variations are not to be regarded as a departure from the spirit and scope of the invention, and all such modifications as would be obvious to one skilled in the art are intended to be included within the scope of the following claims.

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